

॥ सा विद्या या विमुक्तये ॥



स्वामी रामानंद तीर्थ मराठवाडा विद्यापीठ, नांदेड

“ज्ञानतीर्थ” परिसर, विष्णुपुरी, नांदेड - ४३१६०६ (महाराष्ट्र)

SWAMI RAMANAND TEERTH MARATHWADA UNIVERSITY NANDED

“Dnyanteerth”, Vishnupuri, Nanded - 431606 Maharashtra State (INDIA)

Established on 17th September 1994 – Recognized by the UGC U/s 2(f) and 12(B), NAAC Re-accredited with 'A' Grade



ACADEMIC (1-BOARD OF STUDIES) SECTION

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संलग्नित महाविद्यालयांतील विज्ञान व तंत्रज्ञान विद्याशाखेतील पदवी स्तरावरील प्रथम वर्षाचे CBCS Pattern नुसारचे अभ्यासक्रम शैक्षणिक वर्ष २०१९-२० पासून लागू करण्याबाबत.

प रि प त्र क

या परिपत्रकान्वये सर्व संबंधितांना कळविण्यात येते की, दिनांक ०८ जून २०१९ रोजी संपन्न झालेल्या ४४व्या मा. विद्या परिषद बैठकीतील ऐनवेळचा विषय क्र.११/४४-२०१९ च्या ठरावानुसार प्रस्तुत विद्यापीठाच्या संलग्नित महाविद्यालयांतील विज्ञान व तंत्रज्ञान विद्याशाखेतील पदवी स्तरावरील प्रथम वर्षाचे खालील विषयांचे C.B.C.S. (Choice Based Credit System) Pattern नुसारचे अभ्यासक्रम शैक्षणिक वर्ष २०१९-२० पासून लागू करण्यात येत आहेत.

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|---|---------------------------------------|
| 1. Agricultural Microbiology | 18. Dyes and Drugs |
| 2. Agrochemicals & Fertilizers | 19. Electronics |
| 3. Analytical Chemistry | 20. Environmental Science |
| 4. B.C.A. | 21. Fishery Science |
| 5. B.Voc. (Food Processing, Preservation and Storage) | 22. Food Science |
| 6. B.Voc. (Web Printing Technology) | 23. Geology |
| 7. Biochemistry | 24. Horticulture |
| 8. Bioinformatics | 25. Industrial Chemistry |
| 9. Biophysics | 26. Information Technology (Optional) |
| 10. Biotechnology (Vocational) | 27. Mathematics |
| 11. Biotechnonology | 28. Microbiology |
| 12. Botany | 29. Network Technology |
| 13. Chemistry | 30. Physics |
| 14. Computer Application (Optional) | 31. Software Engineering |
| 15. Computer Science (Optional) | 32. Statistics |
| 16. Computer Science | 33. Zoology |
| 17. Dairy Science | |

सदरील परिपत्रक व अभ्यासक्रम प्रस्तुत विद्यापीठाच्या www.srtmun.ac.in या संकेतस्थळावर उपलब्ध आहेत. तरी सदरील बाब ही सर्व संबंधितांच्या निदर्शनास आणून द्यावी.

‘ज्ञानतीर्थ’ परिसर,
विष्णुपुरी, नांदेड - ४३१ ६०६.
जा.क्र.: शैक्षणिक-०१/परिपत्रक/पदवी-सीबीसीएस अभ्यासक्रम/
२०१९-२०/२९२

दिनांक : ०३.०७.२०१९.

प्रत माहिती व पुढील कार्यवाहीस्तव :

- १) मा. कुलसचिव यांचे कार्यालय, प्रस्तुत विद्यापीठ.
- २) मा. संचालक, परीक्षा व मूल्यमापन मंडळ यांचे कार्यालय, प्रस्तुत विद्यापीठ.
- ३) प्राचार्य, सर्व संबंधित संलग्नित महाविद्यालये, प्रस्तुत विद्यापीठ.
- ४) साहाय्यक कुलसचिव, पदव्युत्तर विभाग, प्रस्तुत विद्यापीठ.
- ५) उपकुलसचिव, पात्रता विभाग, प्रस्तुत विद्यापीठ.
- ६) सिस्टम एक्सपर्ट, शैक्षणिक विभाग, प्रस्तुत विद्यापीठ.

स्वाक्षरित / -

उपकुलसचिव

शैक्षणिक (१-अभ्यासमंडळ) विभाग

Swami Ramanand Teerth Marathwada University, Nanded

Faculty of Science

B.Sc. I (First) Year (Semester I & II)

Analytical Chemistry

CBCS

Course Structure, Semester-I & II (w.e.f.2019-20)



Faculty of Science

B.Sc. First Year

ANALYTICAL CHEMISTRY

Sr. No.	Semester	Paper No.	Name of the Course	Instruction Hrs/ week	Total period	Internal Evaluatio	Marks of Semester	Total Marks	Credits
1	I	CCAC I (Section A)	General Concepts of Analytical Chemistry-I (PI)	03	45	10	40	50	2
		CCAC I (Section B)	Basic Analytical Chemistry-I (PII)	03	45	10	40	50	2
2	II	CCAC II (Section A)	General Concepts of Analytical Chemistry-III (PIII))	03	45	10	40	50	2
		CCAC II (Section B)	Basic Analytical Chemistry-IV [PIV]	03	45	10	40	50	2
		CCACP I [CCAC I & II (Section A & B)]	Practical's based on Section A & Section B of CCAC I & CCAC II (PV)	04	20 Practical	20	80	100	4
									12

Total Credit: 12

The syllabus is based on six (3x2) theory periods and 4 practical periods per batch per week. Candidates should require passing separately in theory and practical examination.

Marks distribution:

- 1) Theory exam: 40 marks (30+10 for each paper)
- 2) Internal evaluation: 10 marks (Test or Assignment & attendance)



Faculty of Science
B.Sc. Second Year
ANALYTICAL CHEMISTRY

III	CCAC III (Section A)	Inorganic and Organic Analysis – I (P-VI)	03	45	10	40	50	2
	CCAC III (Section B)	Instrumental Methods of Chemical Analysis – I (P-VII)	03	45	10	40	50	2
	CCACP II [CCAC III & IV (Section A)]	Practical's based on P-VI & P-VIII (P-X)	04	20 Practicals	10	40	50	2
	CCACP II [CCAC III & IV (Section A)]	SEC I (1 Skill/ optional)			15×3 = 45	-	-	(02)*
IV	CCAC IV (Section A)	Inorganic and Organic Analysis – II (P-VIII)	03	45	10	40	50	2
	CCAC IV (Section B)	Instrumental Methods of Chemical Analysis – II (P-IX)	03	45	10	40	50	2
	CCAC P III [CCAC III & IV (Section B)]	Practical's based on P-VII & P-IX (P-XI)	04	20 practicals	10	40	50	2
	CCACP III [CCAC III & IV (Section B)]	SEC II (1 Skill / optional)			15×3 = 45	-	-	(02)*
Total credits semester III and IV								12(04)*



Faculty of Science

B.Sc. Third Year

ANALYTICAL CHEMISTRY

ANALYTICAL

CHEMISTRY

V	DECAC I (Section A)	Modern Techniques Of Chemical Analysis - I (P-XII)	03	45	10	40	50	2
	DECAC I [(Section B) Elective]	Applied Analytical Chemistry - I (P- XIII)	03	45	10	40	50	2
	DECAC PI [DECAC I & II (Section A)]	Practical's based on P- XII & P XIV(P-XVI)	04	20 Practicals	10	40	50	2
	DECACP II [DECAC I & II (Section A)]	SEC III (1 Skill/ optional)			15×3 = 45	-	-	(02)*
VI	DECAC II (Section A)	Modern Techniques Of Chemical Analysis - II (P-XIV)	03	45	10	40	50	2
	DECAC II [(Section B) Elective]	Applied Analytical Chemistry -II (P-XV)	03	45	10	40	50	2
	DECACP III [DECAC I & II (Section B)]	Practical's based on P- XIII & P-XIV (P-XVII)	04	20 Practicals	10	40	50	2
	DECACP IV DECAC I & II (Section B)	SEC IV (Project))			50	-	50	(2)*
Total credits semester V and VI								12(04)*



Swami Ramanand Teerth Marathwada University, Nanded

Faculty of Science

B.Sc. I (First) Year; Semester - I (w.e.f.2019-20)

Analytical Chemistry; Paper - I

General Concepts of Analytical Chemistry-I

Paper Code – CCAC I (Section A)

Periods: 45 per semester; 03 per week

Credit-2

Unit-I

Scope and Importance of Analytical Chemistry:

10 Periods

Introduction to analytical chemistry, Role of analytical chemistry in sciences, Qualitative analysis, Quantitative analysis; major, minor and trace constituents. Quantitative methods of analysis- classification of analytical methods according to property, parameter measured, size of the sample with explanation. Steps in typical quantitative analysis. Types of analysis complete analysis, partial analysis and assay of ingredients, the analytical chemist and analyst.

Unit-II

Preliminary Operations in Quantitative Analysis:

12 Periods

Introduction, sampling: definitions, purpose of sampling, theory of sampling, types of sampling, sampling of solids, liquids and gases. Preparation of laboratory samples: crushing and grinding of laboratory samples; moisture in samples and drying, determination of water in sample, decomposition and dissolution of samples, some general considerations. Acid treatment, decomposition by flux treatment, decomposition of organic matter (Organic elemental analysis and preparation of solution of sample.

Unit- III

Mole Concept and Concentration Units:

13 Periods

Mole Concept, molecular weight, formula weight, and equivalent weight. compounds) for Concentration units: Molarity, Formality, Normality, Molality, Mole fraction, Percent by weight, Percent by volume, Parts per thousand, Parts per million, Parts per billion, pX, pH, pOH, pM, milliequivalents, Milli moles and Titer, Numerical.

Unit-IV

Aspects of Co-ordination Compounds in Chemical Analysis:

10 Periods

Definition of terms: Co-ordination complex, Co-ordination number, Chelate: difference between complex and chelates. Types of chelating agents, significant properties of metal ions and ligands which influence co-ordination. Stability and stability constant of complexes, Stepwise formation constant. Evidence for complex formation. Application of Complexes in Identification of Metal ions, Separation of metal ions and estimation of metal ions.

Objective(s)	To understand the scope and importance of analytical
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	chemistry, acquire the knowledge about preliminary operation in quantitative analysis, mole concept, concentration units, and aspect of coordination compound in chemical analysis.
Course Outcome(s)	
CO1	Understand the scope and importance of analytical chemistry
CO2	Learn about the preliminary operation carried out in quantitative analysis.
CO3	Learn about mole concept and concentration unit.
CO4	Learn about the aspect and uses of coordination compound in chemical analysis



Faculty of Science
B.Sc. I (First) Year; Semester - I (w.e.f.2019-20)
Analytical Chemistry; Paper - II
 Basic Analytical Chemistry-I
Paper Code – CCAC I (Section B)

Periods: 45 per semester; 03 per week

Credit-2

Unit-I

Measurement of Mass :

14 Periods

Distinction between mass and weight; Types of analytical balances: Sem -micro analytical balance, single pan analytical balance, electronic analytical balance. Two-pan balance (Equal arm balance). General features, principle of construction, working of typical analytical balance. Single pan balance: principle, construction and working, instruction for use of single pan balance. Electronic analytical balance: Principle, construction and working. Sources of errors in weighing and their elimination. Weights and calibration of weights.

Unit-II

Measurement of Volume :

7 Periods

Units of volume, effect of temperature on volume measurement. Apparatus for precise measurement of volume; pipette, burette and volumetric flask & their calibration.

Unit – III

Principles of Volumetric Analysis-I:

12 Periods

Definition of terms: Titrant, titrand, analyte, end point and equivalence point, indicator, standard titrant, titration. Acid-base titration: Theory of acid base indicators, Theory of acid-base titration, titration of strong acid-strong base, weak acid-weak base, strong acid-weak base with titration curve and choice of indicators.

Unit-IV

Principles of Volumetric Analysis-II:

12 Periods

Redox Titration: Theoretical basis of volumetric analysis involving (i) Potassium

Permanganate (ii) Potassium dichromate and (iii) Iodine. **Precipitation titration :** Titration curve for precipitation reaction, end point detection, Mohr's method and Volhard's method. **Complexometric Titration:** Theory of complexometric titration, indicators for EDTA titration, Types of EDTA titration-direct and back titration.

Objective(s)	To acquire the basic knowledge about mass measurement of mass, volume measurement, principles of volumetric analysis.
Course Outcome(s)	
CO1	Learn basic concept about measurement of mass and their instruments.

CO2	Learn about the measurement of volume.
CO3	Get the knowledge about the principles of volumetric analysis.
CO4	Learn about the different types of titrations.



Swami Ramanand Teerth Marathwada University, Nanded

Faculty of Science

B.Sc. I (First) Year; Semester - I (w.e.f.2019-20)

Analytical Chemistry; Paper - III

General Concepts of Analytical Chemistry-II

Paper Code – CCAC II (Section A)

Periods: 45 per semester; 03 per week

Credit-2

Unit-I

Errors in Chemical Analysis :

10 Periods

Replicate analysis, reliability of analytical data, mean and median & range precision and accuracy, methods of expressing precision and accuracy: deviation, mean deviation, relative mean deviation, and standard deviation. Errors, absolute error, relative error. Determinate errors, classification of determinate errors and their minimization, indeterminate error and normal frequency distribution curve.

Unit-II

Statistical Treatment of Analytical Data:

13 Periods

Statistical treatment of analytical data, confidence limits, students T-test, rejection of data: Q-test, 4d rule and 2.5d rule. Graphical representation of results, methods of averages, methods of least squares. Significant figures, Reporting of analytical data, Numericals.

Unit-III

Introduction to Chromatographic Techniques:

10 Periods

Introduction, general principle of chromatography, classification of chromatographic techniques. Principle, technique and applications of paper and thin layer chromatographic techniques.

Unit-IV

Purification Methods used in Organic Chemistry:

12 Periods

Theory of distillation, fractional distillation & Crystallization.

Objective(s)	To acquire the basic knowledge about errors in chemical analysis, statistical treatment of analytical data. Aware about separation and purification technique such as chromatography and distillation, crystallization.
Course Outcome(s)	
CO1	Understand about the errors in chemical analysis.
CO2	Learn about the statistical treatment of analytical data.
CO3	Get the basic knowledge about chromatography techniques.

CO4	Learn about the purification methods like distillation and crystallization.
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Swami Ramanand Teerth Marathwada University, Nanded

Faculty of Science

B.Sc. I (First) Year; Semester - I (w.e.f.2019-20)

Analytical Chemistry; Paper - IV

Basic Analytical Chemistry-II

Paper Code – CCAC II (Section B)

Periods: 45 per semester; 03 per week

Credit-2

Unit-I

Gravimetric Methods of Analysis-I :

12 Periods

Introduction to gravimetric analysis, general principle, entire gravimetric procedure and gravimetric steps. Gravimetric conversion Factor (GCF) - illustrations with reference to sulfate, chloride, ferric, calcium and phosphate as analyte ions.

Precipitation: Saturation, super saturation, nucleation and crystal growth. Properties of precipitates-particle size, colloidal state; types of precipitates-crystalline, curdy and gelatinous precipitates.

Unit-II

Gravimetric Methods of Analysis-II :

12 Periods

Purity of precipitates, co-precipitation, post-precipitation and procedures to minimize. Factors affecting precipitation. Precipitation from homogeneous solution. Ageing precipitate, filter papers, filter mats, Gooch crucible, Sintered glass crucible, washing drying and ignition of precipitate. Comparison of gravimetric analysis with volumetric analysis.

Unit – III

Types of Precipitants and Their Applications :

11 Periods

Inorganic precipitants, organic precipitants, their advantages and disadvantages. Uses of inorganic precipitants: silver nitrate for Chloride, dilute sulfuric acid for barium and lead, barium chloride for sulfate and ammonium hydroxide for iron (III). Uses of organic precipitants: dimethyl glyoxime for Nickel, 8-hydroxy quinoline for aluminum and α -benzoinoxime (Cupron) for copper.

Unit-IV

Solvents and Reagents:

10 Periods

Solvents: Solute, Solvent & Solution, classification of solvents (i) Protic and aprotic (ii) Acidic, basic amphiprotic and neutral (iii) Aqueous and non-aqueous (iv) Polar and non-polar. Each type to be explained with at least one example.

Reagents: Classification of reagents according to their action; (i) acids (ii) bases (iii) salts (iv) complexing agents (v) oxidizing and reducing agents (vi) precipitating agents (vii) chelating agents. Each type to be explained with at least one suitable example. Primary and secondary standards: Definition, characteristics, uses, examples for different types of reactions.

Objective(s)	To acquire the basic knowledge about gravimetric analysis, learn about the precipitant and their application, and learn about the solvent and reagent in chemical
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	analysis.
Course Outcome(s)	
CO1	Learn basic concept about methods of gravimetric analysis and precipitation process.
CO2	Familiar and learn the different types of precipitation process.
CO3	Get the knowledge about the precipitants and their application in chemical analysis.
CO4	Learn about the different types of solvent and reagent used in gravimetric analysis.



Swami Ramanand Teerth Marathwada University, Nanded

Faculty of Science

B.Sc. I (First) Year; Semester - I & II (w.e.f.2019-20)

Paper Code – CCACP I [CCAC I & II (Section A & B)]

Analytical Chemistry; Paper - V

Laboratory Course-I

Periods: 45 per year; 04 per week

Credit-4

Note: Out of 23 experiments 18 experiment should be completed.

1. Calibration of volumetric apparatus: Pipette / Standard flask.
2. Preparation of standard solution of potassium hydrogen phthalate and standardization of sodium hydroxide solution.
3. Preparation of standard solution of $K_2Cr_2O_7$ and standardization of given $FeSO_4$ solution.
4. Preparation of standard solution of $(COONa)_2$ and standardization of given $KMnO_4$ solution.
5. Preparation of $Na_2S_2O_3$ solution and its standardization using standard $K_2Cr_2O_7/ KIO_3$ solution.
6. Preparation of standard solution of $NaCl$ and standardization of given $AgNO_3$ solution.
7. Separation of metal ions (Cu^{2+} , Pb^{2+} and Cd^{2+}) / (Z^{2+} , Co^{2+} & Ni^{2+}) by paper chromatography.
8. Assay of commercial sodium hydroxide/ barium hydroxide.
9. Assay of H_2O_2 solution.
10. Assay of formaldehyde.
11. Determination of alkalinity of water sample.
12. Determination of free chloride in a sample of water.
13. Determination of acetic acid content in a commercial sample of vinegar.
14. Determination of moisture content in a soil/ coal sample.
15. Estimation of HCl and CH_3COOH in mixture using acid base indicators.
16. Estimation of iodine in the given solution using standard $Na_2S_2O_3$ solution.
17. Preparation of $EDTA$ solution and its standardization using standard Zn^{++} solution.
18. Estimation of Al^{+++} in the given solution using standard $EDTA$ solution (Back Titration).
19. Estimation of calcium in the given sample of Lime stone or Dolomite or Calcite using standard $EDTA$ solution.
20. Estimation of ester by hydrolysis.
21. Determination of Carbon Dioxide in polluted water sample.
22. Determination of Calcium in Calcium Gluconate.

Objective(s)	To make student aware about the calibration of
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	glassware as well as handle the laboratory instruments. To trend the student carrying out various titrations and estimations.
Course Outcome(s)	
	Learn basic ideas about the calibration of glassware and handling of glassware and laboratory instruments. Students are get trend in carry out different titration and estimations.



Swami Ramanand Teerth Marathwada University, Nanded
Faculty of Science
Analytical Chemistry

Reference Books:

1. Analytical chemistry: an introduction: D. A. Skoog, D. M. West and F. J. Holler, Saunders College publishers, 6th edition.
2. An introduction to analytical chemistry, S. A. Iqbal, M. Satake, Y. Mido and M. S. Shethi.
3. College analytical chemistry: Joshi, Baliga and Shetty, Himalaya Publishing house.
4. Qualitative analysis: Day and Underwood.
5. Qualitative inorganic analysis: A. I. Vogel.
6. Principles of analytical chemistry: Pandit and Soman.
7. Analytical chemistry, G. D. Christian, J. Wiley eastern press Ltd.
8. Analytical chemistry: Alka Gupta.
9. Basic concepts of analytical chemistry: S. M. Khopkar.
10. Advanced practical organic chemistry: Vishnoi.
11. Qualitative analysis: A laboratory manual: Day and Underwood.
12. Fundamentals of analytical chemistry: D. A. Skoog, D.M. West and H. J. Holler, 7th edition.
13. Analytical chemistry principles: J. H. Kennedy, W. B. S. Saunders pub. Ltd.
14. Analytical chemistry: Principles and Techniques: L. G. Hargis, Prentice Hall.
15. Principles in semi-micro qualitative analysis: G. R. Chatwal edited by M. Arora.
16. Experiments in chemistry: D. V. Jahagirdar.
17. A text book of experimental and calculation in engineering chemistry: S. S. Dara.
18. Analytical chemistry: Pitrzyk and Frank, second edition.
19. Modern analytical chemistry: W. F. Pickering, Marcel Decker INC. New York.
20. Introduction to chromatography: Srivastava and Srivastava.
21. University Practical Chemistry by PC Kamboj, Vishal Publishing Company, Jalandhar.
22. Practical Chemistry (for B.Sc.I, II & III Year Students of All Indian Universities) Dr.O.P.Panday, D.N. Bajpai & Dr. S. Giri, S.Chand & Company, New Delhi.
